

## Effect of reactor configurations on the Suzuki cross-coupling reaction using a carrageenan-based $\text{RhCl}_3$ catalyst

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**Abstract:**  $\text{RhCl}_3$  was heterogenized into an iota-carrageenan polysaccharide support, and the effect of various reactor configurations on catalyst performance in a Suzuki cross-coupling reaction was studied. It was found that performing the reaction in a high-volume, well-agitated, mechanically mixed high volume reactor, or circulating the reaction mixture through the catalyst, which was placed in a tubular reactor, yielded higher product yields in comparison to magnetically stirring or shaking. In addition, the catalyst was also successfully recycled in all the systems. Moreover, the reaction solution in each reactor configuration did not contain any traces of rhodium, while elemental composition of the iota-based rhodium catalyst, *i*- $\text{RhCl}_3$ , following the first and the second cycle in the glass tubular reactor was similar to the fresh *i*- $\text{RhCl}_3$  catalyst, demonstrating the high xerogel stability.

**Keywords:** Heterogeneous catalysis; polysaccharides; rhodium; Suzuki cross-coupling. ©2021 ACG Publication. All right reserved.

### 1. Introduction

Transition metal complexes (TMCs) are very active, selective and versatile catalysts for organic syntheses.<sup>1,2</sup> Yet their separation from the reaction mixture and recycling is usually difficult to accomplish. Thus, despite their evident advantages, TMCs have found limited use to date in large-scale industrial applications.

TMC heterogenization is a well-known and extensively studied method, yielding a heterogeneous analog to the homogenous catalyst.<sup>3-5</sup> An optimal heterogenized catalyst ought to have all the advantages of homogeneous catalysts, including high performance, simple separation and reusability. A variety of inorganics and organics supports have been employed for TMC heterogenization. These include renewable and biodegradable polysaccharides produced by various living organisms, which have found many applications nowadays.<sup>6-10</sup>

Among the variety of heterogeneous systems, palladium-based polysaccharide heterogeneous catalysts have been widely investigated,<sup>11-14</sup> and most of these new palladium catalysts have been tested in the Suzuki cross-coupling of halobenzenes and organoborons.<sup>15-20</sup> Yet, over the years, other metals, such as rhodium,<sup>21,22</sup> platinum,<sup>23</sup> nickel,<sup>24,25</sup> and copper,<sup>26</sup> have also been tested in a Suzuki cross-coupling, and their applications within heterogeneous catalytic systems have thereby also been investigated.

We recently synthesized several heterogeneous catalysts by immobilizing palladium and rhodium salts and complexes in polysaccharide matrixes, yielding 'sponge-like' xerogels.<sup>27-31</sup> These

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catalysts were successfully used, separated and recycled in the Suzuki cross-coupling reaction. Furthermore, rhodium chloride was immobilized on iota-Carrageenan (*i*) polysaccharide without the addition of a ligand i.e., *i*-RhCl<sub>3</sub>.<sup>31</sup> In addition, it was demonstrated by XPS analysis that the rhodium oxidation state, Rh<sup>+3</sup>, was not changed during the immobilization or the reaction, and that the atomic ratio of Rh:Cl in the heterogeneous catalyst was around 3:1. Furthermore, the rhodium was immobilized to the *i* via the sulfate groups on its backbone without any ligand. Thus it offers very simple, cheap and more environmentally friendly system, in comparison to palladium.

We report in this paper on the use of *i*-RhCl<sub>3</sub> as a heterogeneous catalyst in Suzuki cross-coupling of iodobenzene and phenylboronic acid, using various high-volume laboratory reactor configurations.

## 2. Experimental

### 2.1. Reagents and Polysaccharide Supports

All chemicals (analytical grades) and *i* polysaccharide were purchased from Sigma-Aldrich Israel.

### 2.2. Reagents and Polysaccharide Supports

The heterogeneous catalyst was prepared as described before.<sup>30</sup> Briefly, 1 g of *i* was dissolved in 100 mL double-distilled water until a homogeneous solution was obtained, and was added to a vial with 10 μmol or 20 μmol RhCl<sub>3</sub> in 3 mL DDW and mixed by vortexing for homogenization. It was then poured into a 15 mL polypropylene tube, sealed, mixed, and deep-frozen at -20 °C for 24 h, until all the liquid was frozen. Finally, the seal was removed and the tube was covered with a paraffin sheet pierced by a disposable toothpick and lyophilized for 48 h, yielding a 'sponge-like' xerogel that was cut into ~ 1cm×1cm pieces and added to the reaction mixture.

### 2.3. Reaction Procedure

In a typical procedure, 20 μmol of RhCl<sub>3</sub> (homogenous or heterogeneous in the form of 1 xerogel with 20 μmol catalyst, i.e., 20 μmol *i*-RhCl<sub>3</sub> or 2 xerogels with 10 μmol catalyst, i.e., 2×10 μmol *i*-RhCl<sub>3</sub>) were added to a glass vessel, e.g., a 10-250 mL beaker or 250 mL Erlenmeyer, or a 1 L Buchi glass reactor, with 5-200 mL ethanol together with 0.5 g iodobenzene, 0.45 g phenylboronic acid and 0.34 g Na<sub>2</sub>CO<sub>3</sub>. In the beaker system, an up to 200 mL reaction mixture was placed in a preheated oil bath at 50 °C and magnetically stirred at 100 rpm for 24 h. Alternatively, a 250 mL Erlenmeyer with a 200 mL reaction mixture was placed in a Julabo SW22 shaking water bath, and shaken at 100 rpm and 50 °C for 24 h. In the third system, a 1 L Buchi glass reactor (diameter=164 mm) with a propeller stirrer (diameter=60 mm) was filled with a 200 mL reaction mixture and kept at 100 rpm and 50 °C for 24 h.

The organic phase was analyzed during the reaction to determine the reaction product yield by gas chromatography (GC) by using a HP-5 column. The average sum-square peak ratios were calculated according to a calibrated reactant–product curve. At the end of the reaction, the reaction mixture was cooled to room temperature.

Another system was built of a 2.5 mm×20 mm tabular glass reactor connected by Teflon pipes to a 250 mL glass vessel with two entrances and a Cole Parmer peristaltic pump. The xerogel catalysts (20 μmol *i*-RhCl<sub>3</sub> or 2×10 μmol *i*-RhCl<sub>3</sub> or 2×20 μmol of *i*-RhCl<sub>3</sub>), was situated in the middle of a pipe and trapped by glass wool from both sides. The reactants, 0.5 g iodobenzene, 0.45 g phenylboronic acid and 0.34 g Na<sub>2</sub>CO<sub>3</sub>, were added together with 200 ml ethanol to the vessel, which was placed on a heating and mixing plate. Finally, the glass vessel and the glass reactor were heated to 50 °C in Julabo oil baths. The reaction mixture was circulated through the catalyst by the pump and samples were taken from the vessel and analyzed by GC with a HP-5 column during running the reaction.

### 2.4. Leaching Tests and Catalyst Recycling

Leaching of the catalysts was tested by either performing a second reaction after the catalyst was removed from the primary reaction mixture, while adding corresponding amounts of fresh

substrates and sodium carbonate and mixing for 24 h under similar conditions, or using ICP-OES (Arcos, Spectro) analysis for rhodium leftovers in the reaction medium after the first cycle.

Catalyst recycling was performed by filtrating the catalyst after the first cycle and adding the recovered catalyst to a fresh solution with substrates and base. In the tabular reactor, the catalyst recycling was performed by replacing the reaction mixture with a fresh one (containing the reactants and a base).

## 2.6. Analyses

**Surface Analysis by X-ray Photoelectron Spectroscopy (XPS).** XPS measurements were performed by using an X-ray photoelectron spectrometer ESCALAB 250 ultrahigh vacuum ( $1 \times 10^{-9}$  bar) apparatus with an AlK $\alpha$  X-ray source and a monochromator. The spot size of the X-ray was 500  $\mu\text{m}$ , and survey spectra were recorded with a pass energy (PE) of 150 eV while high energy resolution spectra were recorded with a pass energy (PE) of 20 eV. To correct for charging effects, all spectra were calibrated relative to a carbon C 1s peak positioned at 284.8 eV. The XPS results were processed with AVANTGE software.

**Scanning Electron Microscope (SEM analysis).** Scanning electronic microscopy (SEM) was performed with a Phenom XL Scanning Electron Microscope and a backscatter electron detector equipped with an energy-dispersive X-ray spectroscopy detector (Thermo Fisher). Acceleration voltage was 15 kV and sample pressure 0.1 Pa.

**Transmission electron microscope (TEM) Analysis.** High-resolution transmission electron microscopic (HRTEM) micrographs of *i*-RhCl $_3$  system were obtained using an EFI Talos F200C transmission electron microscope (TEM) operated at 200kV at room temperature. The samples were prepared by deposition of a drop of ethanol suspension of the crushed sponge catalyst on a carbon-coated Cu grid.

## 3. Results and Discussion

### 3.1. Chemistry

The investigation began by performing a homogeneous reaction in ethanol, using RhCl $_3$  as the catalyst (Table 1). As biphenyl can be also synthesized via homo-coupling of phenylboronic acid or iodobenzene, (i.e. Ullmann coupling), reference reactions with or without the addition of RhCl $_3$ , using either phenylboronic acid or iodobenzene, were first tested and did not yield any biphenyl. The reactions were first performed under magnetic stirring (MS) in different volumes, from 50 mL to 200 mL (Table 1, entries 1-3), with a constant substrate to catalyst ratio, but with different substrates and catalyst concentrations. As illustrated in Table 1, it was found that changing the reaction volume only slightly changed the product yields.

**Table 1.** Comparison of homogeneous and heterogeneous reactions with *i*-RhCl $_3$  in different systems<sup>1</sup>

Entry	Mode	System	Volume (mL)	Product yield (%)	Product yield (%) <sup>2</sup>	Product yield (%) <sup>3</sup>
1	Homogeneous	MS	50	65	-	-
2	Homogeneous	MS	100	64	-	-
3	Homogeneous	MS	200	62	-	-
4	Heterogeneous	MS	200	18	19	19
5	Homogeneous	SH	200	63	-	-
6	Heterogeneous	SH	200	22	21	19
7	Heterogeneous	SH <sup>4</sup>	200	20	-	-
8	Heterogeneous	MSR	200	48	42	40

<sup>1</sup>Reaction conditions, 0.5 g iodobenzene, 0.45 g phenylboronic acid, 0.34 g Na $_2$ CO $_3$ , 2\*10  $\mu\text{mol}$  *i*-RhCl $_3$ , ethanol, 50  $^\circ\text{C}$ , 24 h.

<sup>2</sup>1<sup>st</sup> reuse.

<sup>3</sup>2<sup>nd</sup> reuse.

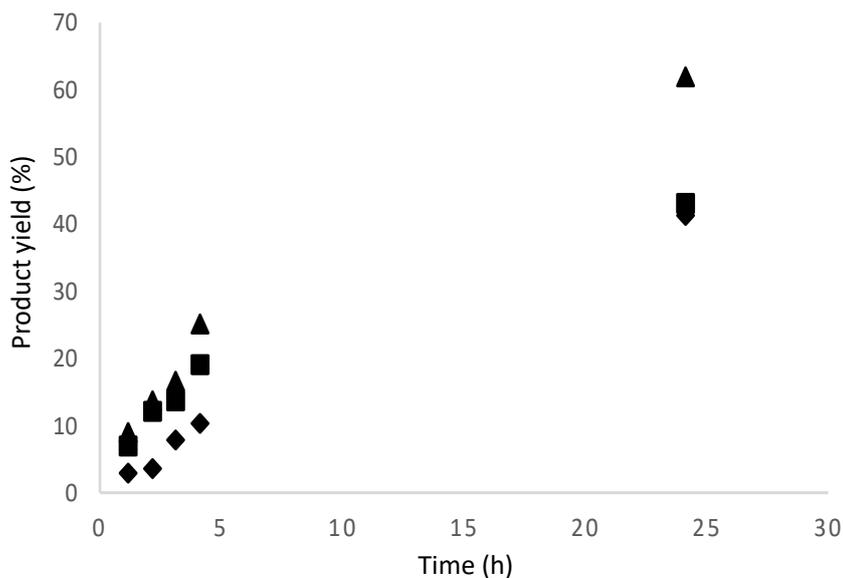
<sup>4</sup>20  $\mu\text{mol}$  *i*-RhCl $_3$ .

As expected,<sup>31</sup> performing the heterogeneous reaction with *i*-RhCl<sub>3</sub> under MS (Table 1, entry 4) yielded lower product yields compared to the homogeneous reaction (Table 1, entry 3), 18% and 62% after 24 h, correspondingly, probably due to mass transfer limitation of the substrates and the product through the support. However, it is worth mentioning that the catalyst was seen to float on top of the solvent, regardless the shape of the vessel, e.g., round glass or Erlenmeyer. Thus, in order to increase the dispersion of the sponge pieces in the solvent, the reaction was also performed in a high volume under shaking conditions (SH). As it turns out, the heterogeneous catalyst was dispersed more effectively in the Erlenmeyer that was shaken, yet the product yield of the heterogeneous system was only slightly higher, reaching 22% (Table 1, entry 6). Moreover, it was still much lower than that of the corresponding homogeneous system (Table 1, entry 5), which yielded a product yield similar to that of the reaction under magnetic stirring (Table 1, entry 3). Furthermore, using only a single sponge with 20 μmol catalysts instead of two sponges with 10 μmol catalyst resulted in only a slightly lower product yield (Table 1, entry 7). This might be attributed to the lower dispersion of the catalyst in the support.

Finally, the Suzuki cross-coupling of iodobenzene and phenylboronic acid was also performed in a mechanically stirred reactor (MSR), equipped with a jacket, where agitating and heat transfer are expected to be much more efficient (Table 1, entry 8). The first step was to find the optimal stirring speed (rpm) that could be applied without crushing the xerogel. It was found that no significant difference was observed between 50 rpm and 100 rpm, whereas a further increase in the stirring speed to 200 rpm led to a vortex and crushed the xerogel so it could not be recycled. Thus, 100 rpm was chosen as the stirring speed. This time, not only was the dispersion of the xerogel in the reaction mixture much greater, the reaction product yield after 24 h was also much higher: 48%, a two-fold increase over magnetic string (Table 1, entry 4) or shaking (Table 1, entry 6). This stresses the importance of efficient momentum, heat and mass transfer and the huge effect that reactor configuration has on catalytic performance.

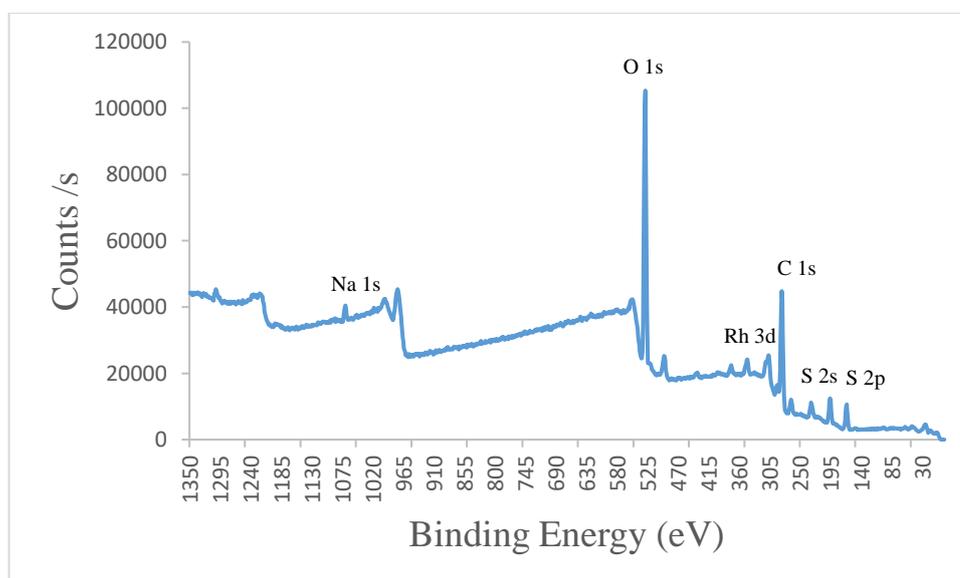
The new heterogeneous catalyst was also separated and reused in the three different systems. It was demonstrated that the product yields at the end of the 2<sup>nd</sup> cycle in each of the configurations (after 48 h) were almost the same as in the 1<sup>st</sup> cycle, while the product yields of the 3<sup>rd</sup> cycle (after 72 h) were slightly lower (Table 1). In all the three systems, the recycling was performed after filtration of the catalyst from the reaction mixture, followed by an addition of the heterogeneous catalyst to a fresh mixture of substrates in ethanol; thus the low decrease in product yield might be attributed to some loss of catalyst during filtration. Yet the recycling results demonstrate that the catalyst did not lose its activity and was well heterogenized.

As the configuration of the reactor has a profound effect on catalyst performance, and continuous reactors are an important class of reactors, we also investigated the Suzuki cross-coupling with the *i*-RhCl<sub>3</sub> catalyst in a glass tabular reactor. Here, the reaction mixture was circulated through the catalysts continuously. Figure 1 illustrates the product yield change with the reaction time, using two sponges of iota with 10 μmol RhCl<sub>3</sub>, one sponge with 20 μmol of RhCl<sub>3</sub>, and 2 sponges with 20 μmol of RhCl<sub>3</sub>. It can be seen that after 24 h the product yields were 41%, 43%, and 62%, correspondingly, much higher than that of the corresponding stirred and shaken systems (Table 1). In addition, as can be seen in Figure 1, again using two sponges with 10 μmol catalyst each, product yields were almost the same as one sponge with 20 μmol of catalyst. Furthermore, employing two sponges with 20 μmol of RhCl<sub>3</sub> each yielded higher product yield of 62% after 24 h, yet the product yield was not doubled. Finally, the reaction mixtures of the various cycles in all the configurations were run for an additional 24 h under similar reaction conditions after removal of the catalyst by filtration. In all the heterogeneous reactions the product yields after removal of the catalyst and mixing for an additional 24 h were unchanged. Furthermore, ICP-OES detection of the reaction mixtures after 24 h reaction and removal of the heterogenous catalysts did not show any rhodium leftovers, as rhodium content was below the detection limit (< 1ppm), which indicates that the catalyst was indeed heterogeneous. Therefore, the recycling results prove that changing the reactor configuration has no effect on the activity of the catalyst after running the reaction in the form of a polysaccharide-based xerogel, thus demonstrating its superior stability.



**Figure 1.** Reaction product yields in a tubular system: (♦)  $2 \times 10 \mu\text{mol } i\text{-RhCl}_3$ , (■)  $20 \mu\text{mol } i\text{-RhCl}_3$ , (▲)  $2 \times 20 \mu\text{mol } i\text{-RhCl}_3$ . Reaction conditions, 0.5 g iodobenzene, 0.45 g phenylboronic acid, 0.34 g  $\text{Na}_2\text{CO}_3$ ,  $2 \times 10 \mu\text{mol } i\text{-RhCl}_3$ , 200 mL ethanol,  $50^\circ\text{C}$

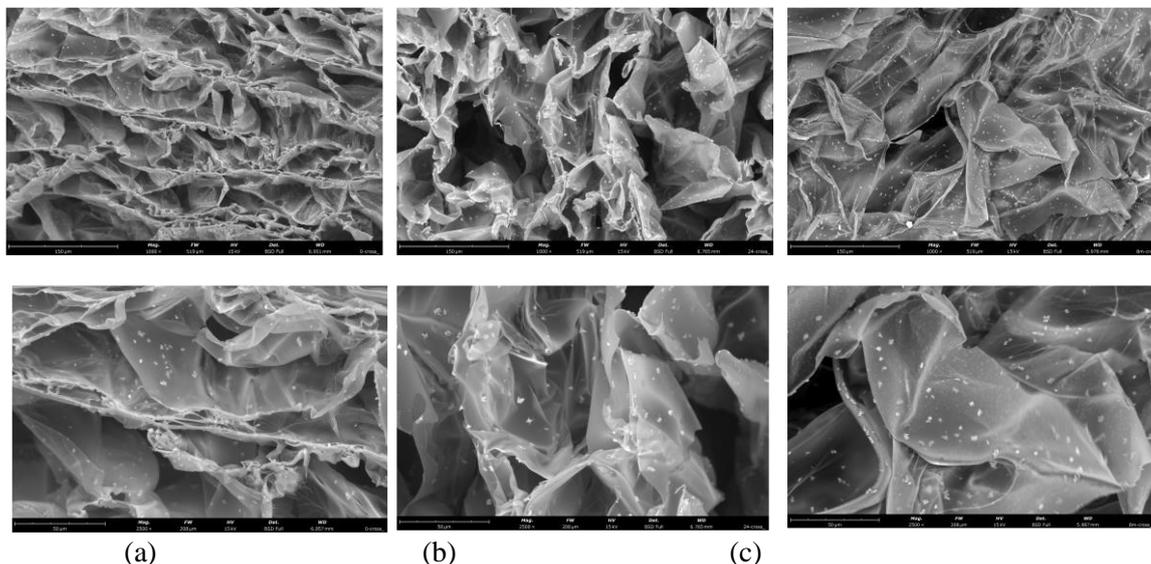
In order to gain more insight into the stability of the heterogeneous catalyst and to evaluate its effectiveness as a reusable catalyst, it was also very important to confirm that the rhodium was embedded in the polysaccharide matrix. Indeed, an XPS survey spectrum of *i*-RhCl<sub>3</sub> confirms that all the elements used to form the 'sponge-like' xerogel catalyst are present (Figure 2). The elemental ratios between the elements was also similar to previously published data<sup>31</sup>. As expected, focusing on the XPS spectrum of *i*-RhCl<sub>3</sub> in the Rh3d region yielded a typical doublet that indicated that the rhodium exists in the oxidation state of +3 (i.e.  $\text{Rh}^{3+}$ ). The atomic percentage of the rhodium measured within the sample was 0.6%, which was in line with our previous publications<sup>31</sup>.



**Figure 2.** XPS survey spectrum for *i*-RhCl<sub>3</sub> before reaction

The next step was to study the changes in the heterogeneous catalyst after the 1<sup>st</sup> and the 2<sup>nd</sup> cycles in the tubular system. Accordingly, the morphology and the elemental composition of *i*-RhCl<sub>3</sub> before the reaction (i.e. fresh *i*-RhCl<sub>3</sub>) and following the 1<sup>st</sup> cycle (after 24 h of reaction conditions) and the 2<sup>nd</sup> cycle (after 48 h), was detected using SEM-EDS. At first, cross-sectional pieces of the different

samples were prepared by cutting the sponge into thin slices. All the slices were characterized by a similar structure, composed of an ordered porous surface (Figure 3). Notably, the exposure of the catalyst to the reaction conditions only slightly affected sponge morphology, as was demonstrated by an increase of void space in the polysaccharide matrix when the catalyst was exposed to reaction conditions for a longer time.



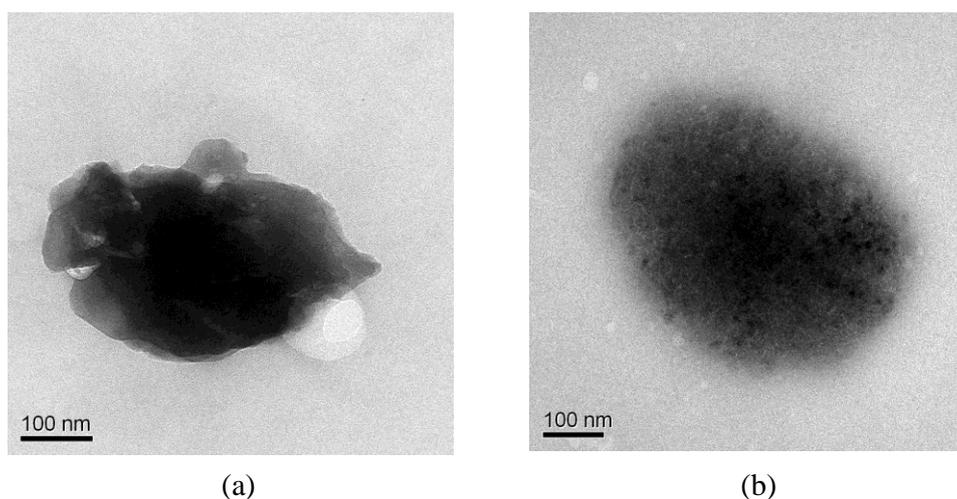
**Figure 3.** SEM analysis of cross-sectional pieces of *i*-RhCl<sub>3</sub> catalyst (a) before reaction; (b) after the 1st cycle (after 24 h) and (c) after the 2nd cycle (after 48 h) in the glass tubular reactor

The composition of the rhodium in the xerogel after the 1<sup>st</sup> and the 2<sup>nd</sup> cycles was similar to the fresh *i*-RhCl<sub>3</sub> catalyst xerogel (~3.00%wt) (Table 2). A slight change in S:Rh ratio was observed following the 2<sup>nd</sup> cycle in comparison to the fresh *i*-RhCl<sub>3</sub> or to the catalyst after the 1<sup>st</sup> cycle. This phenomenon may be attributed to sintering of the rhodium, which could occur due to continuous heating. These results demonstrate that the xerogel did not change its basic compositional structure and exhibited a stabilized structure during the reaction.

**Table 2.** Elemental composition of *i*-RhCl<sub>3</sub> before reaction and following the 1<sup>st</sup> and 2<sup>nd</sup> cycles

Element Symbol	Before Reaction		After the 1 <sup>st</sup> cycle		After the 2 <sup>nd</sup> cycle	
	Atomic %	Weight %	Atomic %	Weight %	Atomic %	Weight %
C	34.7	21.7	32.6	20.3	35.1	23.2
O	42.6	35.5	45.2	37.5	47.6	41.9
Na	1.42	1.70	1.00	1.20	1.13	1.40
S	9.53	15.9	9.50	15.8	7.19	12.7
Cl	2.06	3.80	2.77	5.10	2.51	4.91
K	3.58	7.03	5.56	11.3	3.76	8.10
Ca	5.61	11.7	2.83	5.90	2.18	4.79
Rh	0.50	2.67	0.54	2.90	0.53	3.00

The fresh catalyst and the catalyst that were obtained after 24 h of reaction time in the glass tubular reactor were also analyzed by TEM. The TEM micrograph of *i*-RhCl<sub>3</sub> before the reaction was characterized by an unordered micro-structure, and no rhodium nanoparticles were detected (Figure 4a), while the TEM image of *i*-RhCl<sub>3</sub> after the reaction also yielded a similar structure (Figure 4b), yet with small particles that might be attributed to a change in the polysaccharide support. In addition, the shape of the sponge after the reaction was more symmetric and a bit lighter with spots in different colors, which indicate that the texture of the sponge had changed. These observations suggest that the sponge kept its form under these reaction conditions, indicating that the *i*-RhCl<sub>3</sub> is stable and it is suitable to be used as heterogeneous catalyst.



**Figure 4.** TEM micrographs of (a) fresh *i*-RhCl<sub>3</sub>, (b) *i*-RhCl<sub>3</sub> after reaction

The *i*-RhCl<sub>3</sub> catalyst used in this study was chosen as a representative catalyst to compare the different configurations. To improve the catalytic activity, other metals or TMCs, which have higher activity in a Suzuki cross coupling, can be used. In addition, improving the preparation of the polysaccharide support and increasing the porosity and the surface area, for instance by preparing an aerogel, might increase the product yield.

#### 4. Conclusion

RhCl<sub>3</sub>, immobilized on an iota-carrageenan xerogel by a very simple and straightforward method, was tested in a Suzuki cross-coupling reaction in various reactor configurations. It was found that increasing the solvent volume from 50 ml to 200 ml while using a simple magnetically stirred system and retaining the reactants and the catalyst amounts, only slightly changed the product yields of the homogenous system. Yet the immobilization of the catalyst decreased the product yield due to mass transfer limitations. Shaking the reaction mixture instead of magnetically stirring it yielded better mixing of the heterogeneous system, but almost the same product yield. The best product yields were observed using the mechanically stirred configuration in a Buchi reactor. Furthermore, the catalysts were easily separated from the reaction mixture and recycled, without catalyst loss, while retaining their catalytic performance characteristics. Finally, use of a tubular reactor, in which the reaction mixture was circulated through the xerogel, also yielded high product yields and allowed easy catalyst recycling. Such a setup can resemble the continuous mode systems which have wide current applications in industry.

One of the main obstacles to implementation of new heterogeneous systems at industrial scales is to prepare relatively simple and cheap catalysts that can work in different high-scale reactors. As most of the new polysaccharide-based heterogeneous catalysts that were offered recently in the literature have been examined only at very small scales, it is important to check them also under other configurations. Thus, the significance of these results is considerable for developing new systems for processes meant to be applied at industrial scales. These findings break new ground and call for further investigation. Adapting natural materials for such “green catalysis” can be a substantial boon for industry and lead to far more sustainable methods of production.

#### Supporting Information

Supporting information accompanies this paper on <http://www.acgpubs.org/journal/organic-communications>

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